



Handbook of modern chemistry; inorganic and organic for the use of students

Charles Meymott Tidy

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exposing a solution of the salt to air, a brown liquid results, due to the formation of a normal and basic ferric sulphate, the former remaining in solution and the latter being precipitated ($10\text{FeSO}_4 + \text{O}_2 = 3(\text{Fe}_2\text{O}_3, 3\text{SO}_3) + 2\text{Fe}_2\text{O}_3, \text{SO}_2$). A similar result occurs by the exposure of the solid crystals to air. The aqueous solution of the salt rapidly absorbs nitric oxide. It forms double salts (isomorphous with the magnesian salts), with potassic and ammoniac sulphate. Uses.--In the laboratory it is used as a reducing agent, as, e. g., to precipitate metallic gold from its solutions. In the arts it is used in the manufacture of black dyes and of ink, by the action of tannic acid infusions upon it. (14-) Ferric Sulphate.--Persulphate or sesquisulphate of iron ($\text{Fe}_2\text{S}_2\text{O}_8$), is found native in Chili, as coquimbite ($\text{Fe}_2\text{S}_2\text{O}_8 \cdot 9\text{H}_2\text{O}$). Preparation.--By adding the necessary quantity of H_2SO_4 to FeSO_4 to convert it into $\text{Fe}_2\text{S}_2\text{O}_8$. The mixture is then boiled, the iron being afterwards peroxidized by the cautious addition of nitric acid. Properties.--A non-crystalline, yellow body. (15.) Ferrous Nitrate ($\text{Fe}_2\text{N}_2\text{O}_8 \cdot 6\text{H}_2\text{O}$) is prepared by the action of cold dilute nitric acid on ferrous sulphide; whilst (16.) Ferric Nitrate ($\text{Fe}_2\text{N}_2\text{O}_8 \cdot 12\text{H}_2\text{O}$) is prepared by the action of nitric acid (specific gravity 1.2) on metallic iron. N.B.--Fuming Nitric Acid (specific gravity 1.5) has no action on iron. (See Passive Iron, p. 352.) (17.) Ferrous Carbonate, Protocarbonate of iron (FeCO_3), is found native as spathic iron ore, and as clay iron ore. It is the iron salt commonly found in mineral waters, retained in solution by the carbonic acid dissolved in the water, from which it is precipitated by the action of oxygen, as a red rusty deposit of hydrated ferric oxide. This deposit is fr...

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